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284 Chapter 9 Chemical Reactions Figure 9.3 Science, like all other disciplines, has a specialized language that allows specific information to be communicated in a uniform manner. This reaction between aluminum and bromine can be described by a word equation, a skeleton equation, or a balanced chemical equation.

### **Chapter 9: Chemical Reactions**

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e. molar mass of  $\text{Fe}(\text{NO}_3)_3 = 241.9 \text{ g}$   $272 \text{ g Fe}(\text{NO}_3)_3$   $1 \text{ mol } 241.9 \text{ g} = 1.12 \text{ mol Fe}(\text{NO}_3)_3$  f. molar mass of  $\text{FeSO}_4 = 151.9 \text{ g}$   $12.7 \text{ mg FeSO}_4$   $1 \text{ g } 1000 \text{ mg } 1 \text{ mol } 151.9 \text{ g} = 8.36 \times 10^{-5} \text{ mol FeSO}_4$  g. molar mass of

LiOH = 23.95 g 6.91 103 g LiOH 1 mol 23.95 g = 288.5 = 289 mol LiOH 14. a. molar mass of CaCO<sub>3</sub> = 100.1 g 2.21 10<sup>4</sup> mol CaCO<sub>3</sub> 100.1 g 1 mol = 0 ...

