

chemistry chapter 6 standardized pdf

f. mass of 6 mol C = 6 (12.01 g) = 72.06 g mass of 6 mol H = 6 (1.008 g) = 6.048 g mass of 1 mol O = 16.00 g = 16.00 g molar mass of C₆H₆O = 94.11 g 21. a. molar mass of SO₃ = 80.07 g 49.2 mg SO₃ 1 g 1000 mg 1 mol 80.07 g = 6.14 10⁻⁴ mol SO₃ b. molar mass of PbO₂ = 239.2 g 7.44 x 10⁴ kg PbO₂ 1000 g 1 kg 1 mol 239.2 g = 3.11 10⁵ mol PbO₂

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statements is true? A. An equation is balanced by changing the subscripts as needed. B. Matter is usually, but not always, conserved in a balanced chemical equation.

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F. 1s²2s²2p⁶3s²3p⁶, argon H. 1s²2s²2p⁶3s²3p⁶, krypton G. 1s²2s²2p⁶3s²3p⁶4s², argon J. 1s²2s²2p⁶3s²3p⁸, argon Passage I Use the following passage and table to answer questions 11–13. A student researched the properties of both ionic and covalent compounds. To summarize his findings, he created a table comparing the characteristics of the two ...

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